

Electronic and symmetry engineering of an organic charge-transfer salt with metallic conductivity

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A 1:1 metallic charge-transfer salt is obtained by cosublimation of (Z,E) -(SMe)₂Me₂TTF and TCNQ. X-ray diffraction studies confirm that the formation of a unique assembly comprising donor (in the E -form only) and acceptor molecules in $[(E)$ -(SMe)₂Me₂TTF](TCNQ). The crystal packing features lateral S...S interactions between TTF stacks, which is in sharp contrast to that in (TTF)(TCNQ). Structural analysis and theoretical studies afford a partial charge-transfer ($\rho \approx 0.52$), leading to a system with the electronic structure close to quarter-filled. Resistivity measurements reveal that this material behaves as a metal down to 56 K and 22 K at 1 bar and 14.9 kbar, respectively. The thermopower is negative in the metallic regime, indicating the dominant role of the acceptor stacks for the observed conducting behavior. Analysis of a single-crystal EPR spectra shows the remaining spin susceptibility at 4.3 K, suggesting the importance of the Hubbard U correction. These results highlight the judicious engineering of electronic and geometrical effects on the TTF core to control the physical properties of the organic charge-transfer compound.

Introduction

The discovery of metallic behaviour in crystalline organic molecular materials has stimulated a plethora of different research areas covering chemistry and physics.¹ The most notable example is the charge transfer donor-acceptor compound, (TTF)(TCNQ) (TTF = tetrathiafulvalene; TCNQ = 7,7,8,8-tetracyanoquinodimethane), which was the first organic conductor to present a metallic conductivity in a wide temperature range down to 59 K.² This compound features an anisotropic crystal structure where donor and acceptor molecules form segregated stacks which facilitate a good overlap of respectively HOMOs and LUMOs π -type frontier orbitals along the stack direction to render a 1D electronic structure.³ The accumulated knowledge on such a system emphasizes the critical role of structural chemistry and crystal engineering.⁴ In these molecular systems, a subtle change in chemical structure is known to drastically modify geometrical and electronic structures, which in turn enable to discover new



Scheme 1. Tetrathiafulvalene derivatives with different substituents.

exotic physical phenomena and to underpin physical understanding.⁵ Essentially, the charge transfer salts of the TTF-TCNQ family are found in three different types of arrays mostly depending on the type of substituents in the donor and/or acceptor cores.⁴ However, even within a given type of array the physical behavior is modulated by a complex blend of parameters like charge transfer, elastic properties of the lattice or strength of the electron-electron interactions in the different stacks, which can be subtly altered by chemical or structural factors like the nature of the chalcogen atoms in the donors, the tilt angles with respect to the chain axis, etc. The potential of these two-chain conductors to reveal unexpected features of the quasi-one-dimensional (quasi-1D) metals is far from being fully explored and new types of substituted donors and/or acceptors susceptible to enforce new structural or electronic situations are strongly in demand. In this search we have brought our attention to a TTF donor molecule with two thiomethyl- and two methyl-substituents (Z,E) -(SMe)₂Me₂TTF, (Z,E) -1 (Scheme 1). At variance with classical TTF, TMTTF (tetramethyl-tetrathiafulvalene), TTMTTF (tetrathiomethyl-tetrathiafulvalene) or BEDT-TTF (bis(ethylenedithio)-tetrathiafulvalene) molecules with D_{2h} symmetry, such systems can exist into Z (C_{2v}) and E (C_{2h}) forms, which easily

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†Electronic Supplementary Information (ESI) available: CV, additional structural details, derivative of the resistivity, angle-dependence of ESR signals, schematic view of the inverse domain, details of the band calculation. See DOI: 10.1039/x0xx00000x

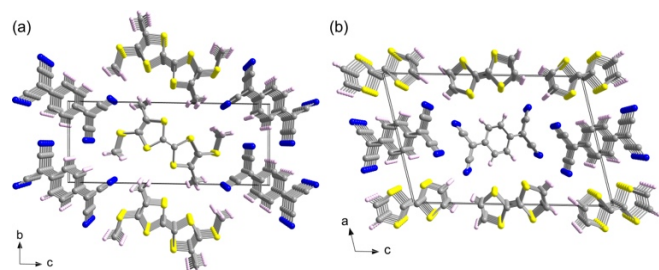


Figure 1. Projection views of the crystal packing in [(*E*)-1](TCNQ) (a) and (TTF)(TCNQ) (b). Yellow, blue, gray, and pink spheres represent S, N, C, and H atoms,

interconvert in solution upon oxidation, making often difficult their (electro)crystallization⁶ despite their readily accessible synthesis. Such a difficulty is well demonstrated by the scarcity of the structurally characterized donor-acceptor two-chain charge transfer salts comprising such *Z*- or *E*-TTF isomers.⁷ Our specific interest for **1** also comes from its first oxidation potential ($E_{1/2} = +0.41$ V vs. SCE, **Figure S1**) close to that of TTF itself (+0.33 V), a consequence of the electron-rich methyl groups which compensate in part the electron withdrawing effect of the thiomethyl groups. We postulate that such a close proximity of electrochemical potentials likely triggers a non-zero charge transfer if this donor is associated with TCNQ,⁸ as demonstrated below.

Results and discussion

Co-crystallization and structural analyses

The neutral (*Z,E*)-**1** was prepared according to the modified literature procedure, by successive deprotection of (*Z,E*)-(SCH₂CH₂CN)₂(Me)₂TTF with CsOH and further alkylation with MeI.⁹ Cosublimation of (*Z,E*)-**1** and TCNQ in a two-zones furnace¹⁰ under vacuum afforded black, flake-shaped crystals of [(*E*)-1](TCNQ) in 3 weeks. Within the co-sublimation tube, we have noticed that the mixture of *Z* and *E* isomers of TTF has been converted to only *E* isomer after the repeated sublimation-recrystallization processes (**Figure S2**). This observation strongly suggests that the co-sublimation method works as a selective crystallization method between different isomers, and therefore, exclusively affords [(*E*)-1](TCNQ) as a co-crystallized product. To the best of our knowledge, such a discrimination of isomers using co-sublimation has never been reported. Note that the solution method for co-crystallization was not successful to obtain single crystals for further analysis. The compound crystallizes in the triclinic system, *P*-1 space group, with both (*E*)-**1** and TCNQ molecules on an inversion centre. The structure consists of segregated stacks of donor and acceptor molecules running both along the crystallographic *a* axis, and alternating along the *c* axis, as shown in Figure 1a. The intermolecular overlaps along the stacks adopt the bond-over-ring motif commonly observed in highly conducting compounds, with plane-to-plane distances of 3.474 and 3.227 Å in the stacks of (*E*)-**1** and TCNQ respectively. Note that those values are larger than those reported in the RT structure of (TTF)(TCNQ) (3.451 and 3.188 Å respectively) (**Table S1**),¹⁰ an indication of possibly weaker

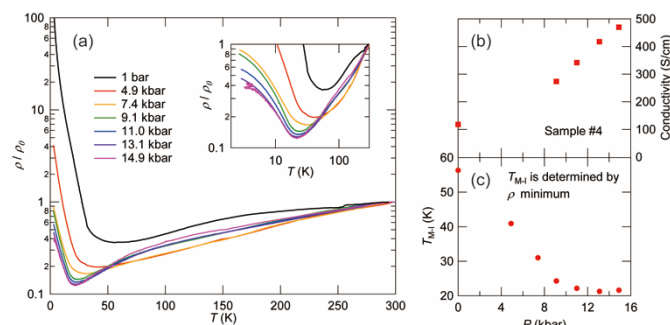


Figure 2. (a) Temperature dependence of the resistivity under pressure. The inset is a logarithmic plot for clearly at low temperatures. Pressure dependence of (b) the room-temperature conductivity and (c) T_{M-I} determined by the resistivity minimum.

intermolecular interactions along the stacks in [(*E*)-1](TCNQ). In the crystallographic *ac* plane (**Figure S3**), all molecules are tilted in the same direction with respect to the (100) plane, with an angle of 22.23 and 31.65° for (*E*)-**1** and TCNQ respectively, leading to an interplanar dihedral angle of 9.42°. A remarkable feature is the relative arrangement of the stacks which allows direct S...S inter-stack contacts along the *b* axis (**Figures 1a and S4**) while such lateral interactions are absent in (TTF)(TCNQ) (**Figure 1b**). Taken together, the arrangement and the packing of donor and acceptor molecules in [(*E*)-1](TCNQ) strongly differs from that of the unsubstituted (TTF)(TCNQ).¹¹ Analogous 1:1 salts were also reported with the *D*_{2h}-symmetry TTFs such as TMTTF and TTMTTF. The (TMTTF)(TCNQ) metallic salt crystallizes in the *P2/c* space group with stacks of TMTTF molecules interacting also sideways, but with large and opposite tilt angles vs. the stacking axis.¹² The (TTMTTF)(TCNQ) charge transfer complex is an insulating compound with mixed stacks in a DDAA sequence, a probable consequence of mismatch of chemical potentials and steric hindrance of the four out-of-plane thiomethyl groups.¹³ These comparisons indicate that the balanced substitution between -SMe and -Me groups in the present donor **1** indeed allows for an original structure which enhances the structural dimensionality of the conducting sheets while avoiding DADA mixed stacks.

To estimate the degree of charge transfer, the bond distances within TCNQ were carefully examined using Kistenmacher's relationship (**Table S2** and **Figure S5**).¹⁴ The calculated charge of TCNQ is 0.52, implying a nearly quarter-filled band system. Note that this value is also coherent with $\rho = 0.59$ estimated from the C≡N stretch of the FT-IR spectrum¹⁵ and the theoretical calculations (*vide-infra*).

Transport properties

The variable-temperature electrical resistivity was measured under different pressures on single crystals along the *a* stacking axis (**Figure 2**). At 300 K, the conductivity is 118 S/cm under ambient pressure, which is comparable to the known organic charge transfer conducting compounds.³ As temperature decreases, the resistivity decreases smoothly to reach a minimum value at 56 K before an abrupt increase, indicating the occurrence of a metal to insulator (M-I) transition. The activation energy in the insulating region is estimated at $E_a = 31$ meV. Before the M-I transition, the

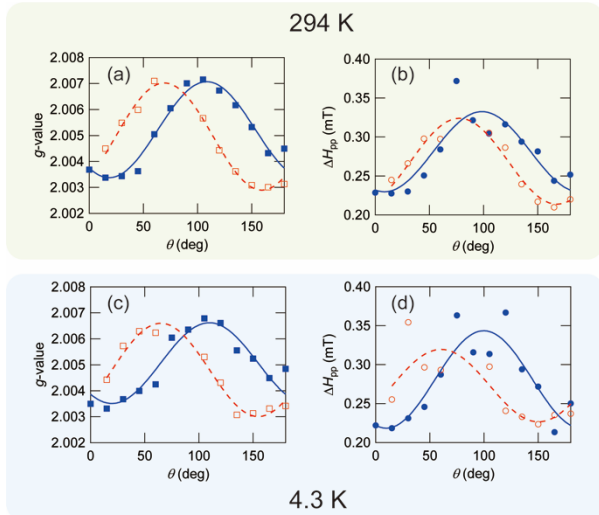


Figure 4. Angle dependence of the g -value and the linewidth at (a)(b) 294 K and (c)(d) 4.3 K, highlighting two independent Lorentzians indicated with red (open) and blue (filled) symbols.

resistivity drop shows a deviation around 130 K, which is confirmed by the maximum of the derivative (Figure S6), likely related to the electronic instability in either the donor or the acceptor stacks (*vide-infra*). Upon increasing pressure, the room-temperature conductivity increases and the M-I transition temperature significantly shifts to lower temperatures, to reach 470 S/cm and 22 K, respectively, at 14.9 kbar (Figures 2b and 2c). At this highest available pressure, the resistivity rise shows no saturation and E_a decreases by half ($E_a = 16$ meV). This is in stark contrast with (TTF)(TCNQ), where the conductivity increases less than twice and the M-I transition temperature hardly varies under similar pressures.^{3a,16} The comparatively high-pressure sensitivity of the present compound is likely attributable to the enhanced structural dimensionality coming from the side-by-side S...S contacts. The applied pressure may indeed reinforce the inter-stack contacts which might ultimately lead to the suppression of the electronic Peierls instability expected for a purely 1D electronic system.

As shown in Figure 3, the thermopower exhibits small negative values over the metallic region, indicating that the transport is dominated by the acceptor stacks.¹⁷ However, in contrast to (TTF)(TCNQ) where it amounts to -40 $\mu\text{V}/\text{K}$ at 300 K, the small absolute value for [(E)-1](TCNQ), -5 $\mu\text{V}/\text{K}$, indicates an increased contribution of the donor stacks to the conductivity. Upon lowering temperature, the thermopower shows a small peak around 150 K then a gradual decrease. Given the negative carriers become more dominant below 150 K, the anomaly at 150 K is probably to be associated with the setting of 4kF instability in the donor stack. Note that such a peak was observed also for (TMTTF)(TCNQ).¹⁷ Below 70 K, the strong upturn of the thermopower is observed, which may be due to the setting of modulation in the TCNQ stacks that finally condense and open a gap around 50 K.

ESR analyses

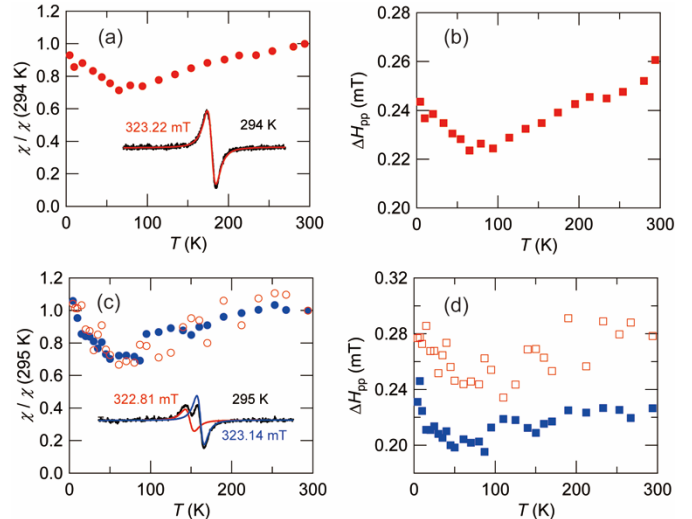


Figure 5. Temperature dependence of the spin susceptibility and the linewidth for (a)(b) the single Lorentzian at 0° and for (c)(d) the double Lorentzians at 30° . The two independent Lorentzians designated with a red (open) symbol and a blue (filled)

Figures 4 shows the angle dependence of the extracted g -values and linewidths at room temperature and at 4.2 K. Although single crystals were chosen carefully, the observed ESR signal was systematically composed of two Lorentzians (Figure S7).¹⁸ Given that the two Lorentzians manifest at the same angles in all three samples investigated, this is likely an inherent property of the compound and the splitting stems from the existence of inversely tilted domains, which have borders in the ab plane (Figure S8). In view of the extremely small deviation from the monoclinic cell ($\beta = 90.292^\circ$), this seems to be inherently associated with the present crystal.

Figure 5 shows temperature dependence of the spin susceptibility and the linewidth estimated from a single Lorentzian at 0° and double Lorentzian at 30° . The spin susceptibility is almost constant with a slight decrease down to 65 K, indicating the Pauli-paramagnetic like behavior. Below 65 K, the spin susceptibility gradually increases. The linewidth also takes a minimum at 65 K and then increases. The observed similar tendency for both Lorentzians supports the same origin of the two signals at 30° . Broadening of the ESR signal is likely related to the nesting of the TCNQ Fermi surface as observed in related compounds, (TTF)(TCNQ) and (TMTTF)(TCNQ).^{18,19} At low temperature, the spin susceptibility remains non-zero even at 4.3 K, indicating charge-spin decoupling in the compound.

Electronic structure calculation

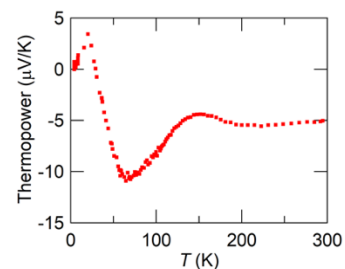


Figure 3. Variable-temperature thermopower (S) of [(E)-1](TCNQ).

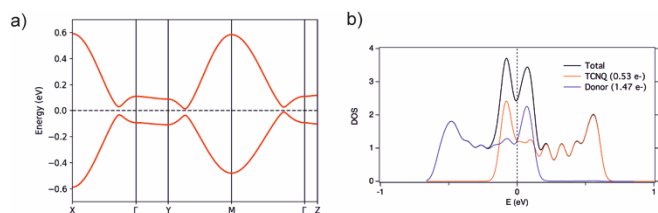


Figure 6. Calculated band structure (a) and Density of states (b) for $\text{Me}_2(\text{SMe})_2\text{TTF-TCNQ}$ at 150 K. The Fermi level is the energy zero. $\Gamma = (0, 0, 0)$, $X = (1/2, 0, 0)$, $Y = (0, 1/2, 0)$, $M = (1/2, 1/2, 0)$ and $Z = (0, 0, 1/2)$ in units of the triclinic reciprocal lattice vectors. The DOS is given in units of electrons per eV and per formula unit.

The electronic structure of $[(E)-1](\text{TCNQ})$ was determined from *first-principles* DFT+U (U being an effective Hubbard correction term²⁰) calculations based on the 150 K crystal structure. Although DFT calculations usually provide a good description of the electronic structure of molecular conductors, two-chain molecular conductors like the present system still face us with some noticeable problems. For instance, the calculated charge transfer in TTF-TCNQ is ~ 0.72 - 0.74 irrespective of the particular functional and basis set used^{21,22} whereas the experimental estimations are 0.55 - 0.59 .⁴ As we have recently discussed,²² this problem mostly concerns two-chain conductors containing TTF or TTF-based donors because of a deficient description of their HOMO bands. These bands are usually found to be wider than the experimental estimations in two-chain systems for which $4k_F$ fluctuations associated with the TTF-based component have been observed. These observations mean that electronic repulsions in the TTF-based chains are important and that they are underestimated by the DFT approach. The simplest way to remedy for this drawback is by using an effective Hubbard correction term for S 3p orbitals which contracts them and leads to narrower HOMO bands and smaller charge transfers.²² Since the present system contains TTF-based chains we have considered the U dependence of the electronic parameters of TTF-TCNQ, HMTTF-TCNQ and TMTTF-TCNQ (see for instance [Figures S9 to S11](#)) and found that a value $U = 6$ eV leads to reasonable results for band widths and charge transfers. For instance, using their room temperature crystal structure the calculated vs. experimental charge transfers (determined at 25 K)⁴ for TTF-TCNQ, TMTTF-TCNQ and HMTTF-TCNQ are $0.57/0.59$, $0.56/0.57$ and $0.68/0.72$. Thus, we have adopted this value in our study.

The calculated band structure and density of states are reported in Figure 6. The upper band at Γ is based on the $(E)-1$ HOMO and the lower band is based on the TCNQ LUMO. The two bands cross along the intrachain $\Gamma-X$ line, at the Fermi energy. However, in these two-chain conductors the avoided crossing between the HOMO and LUMO bands is usually slightly overestimated by DFT so that we believe that the gap along $\Gamma-X$ is smaller and probably irrelevant for most temperatures. The small but non nil dispersion along the interchain b^* -direction ($\Gamma-Y$) of the donor band, which is associated with an $S\cdots S$ contact of 3.767 Å, slightly lowers the energy of the upper band and consequently overlaps with the top of the lower band, definitely leading to a band overlap inside the Brillouin zone and thus, a metallic-type behavior.

The avoided band crossing occurs for a k_F value of $0.13a^*$ and consequently, the charge transfer ($\rho = 4 \times k_F/a^*$) is 0.52 . Integration of the DOS curves (see Figure 5b where the total DOS as well as the separate contributions of the donor and acceptor bands are also shown) leads to $\rho = 0.53$. These values are very similar to the $\rho = 0.52$ one estimated above from the bond lengths of the acceptor thus confirming that the charge transfer must be very close to 0.5 . This situation is clearly different from that of the archetypal (TTF)(TCNQ) for which the charge transfer ($\rho = 0.59$) is definitely far from the quarter-filled/empty case and thus a different physical behavior may be expected. Note that whatever the value of U used in the calculations, the difference with (TTF)(TCNQ) does not practically change ([Figure S9](#)) so that the result is a robust one.

Comparison with (TTF)(TCNQ)

What is the origin of the charge transfer difference with (TTF)(TCNQ)? It may be argued that it is due to the decrease in the HOMO energy induced by the $-SMe$ substituents, illustrated by the increase of first oxidation potential when compared with TTF. However, the charge transfer in these two-chain molecular conductors subtly depends on both electronic and structural factors. A decrease of the charge transfer can also occur if one or both of the HOMO and LUMO bands become narrower. The dispersion of the TCNQ band depends essentially of the $C\cdots C$ contacts along the intrachain a direction. As mentioned above, these contacts in $[(E)-1](\text{TCNQ})$ (Table S1) are clearly longer than even those for the RT structure of (TTF)(TCNQ). Hence we can conclude that the TCNQ band in the present salt must be narrower than in (TTF)(TCNQ). The dispersion of the donor band depends essentially from the six intra-stack $S\cdots S$ contacts. In the RT structure of TTF-TCNQ, the same contacts are again found to be shorter (Table S1), as evidenced also by the cell parameter along the chains direction (3.842 Å for $[(E)-1](\text{TCNQ})$ at 150 K vs 3.819 Å for (TTF)(TCNQ) at RT). However, the $E-1$ donor has two additional sulfur atoms involved in the HOMO. The extra $S_{SMe}\cdots S_{TTF}$ intra-stack contacts increase the HOMO \cdots HOMO interaction with respect to (TTF)(TCNQ) and it is not clear if, in the present compound, the dispersion of the HOMO band should slightly decrease or even increase. Assuming a linear relationship for the thermal evolution of the DFT band widths, we estimate that the widths of the HOMO and LUMO bands for (TTF)(TCNQ) at 150 K should be 0.59 and 0.83 eV, respectively. The calculated band widths for the present compound are 0.70 and 0.68 eV, respectively. As expected, the band width for the TCNQ band decreases in $[(E)-1](\text{TCNQ})$. In contrast, the decrease in the strength of the intrachain $S\cdots S$ interactions of the TTF core is more than compensated by the extra $S_{SMe}\cdots S_{TTF}$ contacts, leading ultimately to an increase of the HOMO band width. However, we note that the variation for the TCNQ band is somewhat larger. In addition, the increase in the donor band width (and thus, of charge transfer) will mostly compensate the intrinsic lowering in the HOMO energy (and thus, the decrease of the charge transfer) due to

the –SMe substituents. We thus suggest that the substantial decrease of the C...C contacts in the TCNQ stacks is the main reason behind the small but significant charge transfer decrease, now very close to the quarter-filled/empty value.

Low temperature behaviour scenario

What are the possible scenario for the low temperature behavior of this compound? An important factor to consider is the filling of the LUMO/HOMO bands which at low temperature may become almost exactly quarter-filled/empty ($\rho = 0.5$) or can keep a slightly different value. In view of the transport measurements reported above and relying on analogies with other TTF-based two-chain salts^{4,17} we suggest the following scenario. The small anomaly occurring around 150 K clearly seen in the thermopower measurements is probably to be associated with the setting of $4k_F$ fluctuations in the donor stacks since the negative carriers become more dominant below 150 K. The strong upturn around 70 K may be due to the setting of a fluctuations regime in the TCNQ stacks which finally condense and open a gap at around 50 K. However, if a charge transfer very near to the quarter-filled/empty case remains at low temperature the modulation on the TCNQ stacks could be a $2k_F$ tetramerization and the regime change in the (*E*)-1 stacks could then be of the $4k_F$ charge-ordering or Mott-dimer types, explaining that the magnetic susceptibility remains non-zero until very low temperatures.²³ X-ray diffuse scattering studies are needed in order to shed light on these issues.

Conclusions

In this study, we have successfully grown a 1:1 co-crystal of (*E*)-1 and TCNQ and have investigated its physical properties in detail. The obtained crystal of [(*E*)-1](TCNQ) features the segregated D and A stacks with the DA packing that is notably different when compared with the analogous charge transfer TCNQ salts comprising the classical TTF donors. In particular, the lateral S...S contacts between donor stacks give rise to the structural two-dimensionality in [(*E*)-1](TCNQ). The band structure of the compound features close to a quarter-filled system with $\rho \approx 0.52$, leading to a metallic behaviour down to 56 K. Moreover, both the conductivity and the metal-to-insulator transition temperature are significantly influenced by the applied pressure, likely due to the enhanced structural dimensionality. In particular, slightly delocalizing the HOMO of the donor by means of the two thiomethyl- and two methyl-substituents leads to a situation where the donor and acceptor band widths are very similar *and* the charge transfer is very near the commensurate value 0.5. To the best of our knowledge the only other two-chain salt of this family exhibiting *both* features is (TMTTF)(DMTCNQ) (DMTCNQ = 2,5-dimethyl-7,7,8,8-tetracyanoquinodimethane).⁴ This salt lies just at the borderline between those where the dominant structural instability changes from $2k_F$ to $4k_F$.⁴ Thus, it would be extremely interesting to study if this is also the case for the present salt. If it is indeed the case, small structural or

electronic modifications could bring substantially different physical behaviors and provide invaluable information to settle the difficult issue of the competition between structural instabilities in two-chain conductors. Work is underway to prove the ground state of [(*E*)-1](TCNQ) and to explore the promising role of (*Z,E*)-1 as electron donor to grow co-crystals with intriguing electronic structures.

Experimental

General methods

The manipulations described below were performed under argon atmosphere with a standard Schlenk technique. Chemicals and materials from commercial sources were used without further purification. Solvents were dried using a commercial solvent purification system from Inert. The (*Z,E*)-(SCH₂CH₂CN)₂(Me)₂TTF was prepared according to the literature procedure.⁹ Melting points were measured on a Kofler hot-stage apparatus. NMR spectra were recorded on a Bruker AV III 300 MHz NMR spectrometer at room temperature using CDCl₃. Column chromatography was performed using Teledyne Isco CombiFlash® Rf400. Cyclic voltammetry measurements were carried out in a one-compartment cell under argon, equipped with a platinum wire, a glassy carbon, and a Saturated Calomel Electrode (SCE) as working, counter, and reference electrodes, respectively, using a Bio-Logic SP-50 potentiationstat. Analyte solutions (10⁻³ M solution of complex) were prepared in 0.1 M solutions of (NBu₄)(PF₆) in CH₂Cl₂.

Synthesis of (*Z,E*)-1

To a solution of the (*Z,E*)-(SCH₂CH₂CN)₂(Me)₂TTF (300 mg, 0.75 mmol) in 75 mL of THF was added under argon a solution of CsOH·H₂O (372 mg, 2.22 mmol) in 10 mL of MeOH dropwise at 0 °C. The resulting solution was stirred at 25 °C for 90 minutes to afford a dark mixture. MeI (0.3 mL, 4.82 mmol) was then added to the flask, and the resulting mixture was stirred for 12 h to give a yellow solution. After quenching the reaction with 60 mL of distilled H₂O, the organic solvent mixture was removed from the reaction under vacuum, and the resulting residue was extracted into CH₂Cl₂ (4 x 60 mL). The organic extracts were dried over magnesium sulfate, filtered, and then concentrated under reduced pressure to give a red oil. The crude oil was purified by flash column chromatography on a silica gel to afford (*Z,E*)-1 as an orange solid (198 mg, 82 %). The product was recrystallized by slowly evaporating the saturated EtOAc solution. Column conditions: silica cartridge 12 g; solid deposit on Celite®; $\lambda_{\text{detection}}$: (254 nm, 280 nm); CH₂Cl₂ in light petroleum with a gradient 0 - 30%. mp = 133 °C. ¹H NMR (300 MHz, Chloroform-*d*) δ = 2.31 (s, 3H), 2.12 (s, 3H).

Crystallization of [(*E*)-1](TCNQ)

Single crystals were prepared by the co-sublimation method¹⁰ in a 30cm vacuum-sealed glass tube (10⁻⁵ mbar) with the temperature for each end set at 130 and 250 °C, for the donor and the acceptor molecules respectively.

X-ray structure determination

Single crystals of (*E*-SMe₂Me₂TTF)(TCNQ) that is suitable for X-ray analysis were coated with Paratone-N oil and mounted on a MicroMounts™ rod. The crystallographic data were collected on a Bruker D8 venture diffractometer equipped with MoK α microsource and PHOTON100 detector at the Centre de Diffraction (CDIFX, Université de Rennes 1, France) at 150 K. Raw data were integrated and corrected for Lorentz and polarization effects using Bruker APEX3 v. 2015.²⁴ The structure was solved by direct methods using the SIR92 program,²⁵ and then refined with full-matrix least-square methods based on F^2 (SHELXL-2014)²⁶ with the aid of the WINGX program.²⁷ All non-hydrogen atoms were refined with anisotropic atomic displacement parameters.

Electrical resistivity and thermopower

Electrical resistivity was measured by the four-probe method down to 2.7 K. Four gold wires (15 μ m diameter) were attached to a crystal using carbon paste, and the resistivity was measured parallel to the crystal long axis. Low-frequency AC current was applied to the outer wires and the resistance was estimated from the voltage measured at the inside wires. Pressure was applied to a crystal by a conventional clamp cell with an oil Daphne #7373 as a pressure medium.²⁸ The absolute value of the pressure was estimated from the resistance change of a manganin wire at room temperature. Results for 9.1 kbar to 14.9 kbar are taken using a different sample from that for 1 bar to 7.4 kbar. For thermopower measurement, a crystal was attached to two copper blocks via gold foils using carbon paste. The copper blocks were alternately heated by glued tip resistors to generate ΔT less than 1K. ΔT was measured by an AuFe-chromel thermocouple. Thermopower was estimated from the slope of the generated voltage ΔV vs. the temperature difference ΔT .²⁹

ESR measurement

ESR measurements were carried out for a single crystal using X-band spectrometer (JEOL JES-TE100). The g -values were calibrated by the 3rd ($g = 2.034$) and 4th ($g = 1.981$) signals of a Mn²⁺ marker. The crystal was located so as to make the long axis parallel to the magnetic field, and rotated by 15° steps up to 180° for the angle dependence measurements (Figure S5). Although the crystal was chosen carefully, the observed signal was composed of two Lorentzians when the crystal long axis was not parallel to the magnetic field. The signal was analyzed by a sum of two Lorentzians,

$$f(H) = \underbrace{-\frac{16I_a \left(\frac{H - H_a^0}{\Delta H_a^{pp}/2} \right)}{\left(3 + \left(\frac{H - H_a^0}{\Delta H_a^{pp}/2} \right)^2 \right)^2}}_{\text{Component 1}} - \underbrace{\frac{16I_b \left(\frac{H - H_b^0}{\Delta H_b^{pp}/2} \right)}{\left(3 + \left(\frac{H - H_b^0}{\Delta H_b^{pp}/2} \right)^2 \right)^2}}_{\text{Component 2}}$$

where H was the magnetic field, I was the intensity of the signal, H^0 was a center of the signal, and ΔH^{pp} was the linewidth of the signal. Temperature dependence measurements were carried out down to 4.3 K for the single Lorentzian and double Lorentzians.

Computational details

The density functional theory (DFT) calculations were performed using a numerical atomic orbitals DFT³⁰ approach developed for efficient calculations in large systems and implemented in the SIESTA code.³¹ The generalized gradient approximation to DFT and, in particular, the functional of Perdew, Burke, and Ernzerhof (PBE)³² has been used. Recently²² we have found that inclusion of a Hubbard correction term²⁰ U for S 3p orbitals is needed for accurately describing the electronic structure of TTF-TCNQ and related two-chain molecular solids for which accurate experimental information on the bandwidth and charge transfer is available. In the present work, we have considered the dependence of the experimentally observed charge transfer for several sulfur-containing two-chain salts and found that a good agreement is found for $U = 6$ eV. Thus, except otherwise stated we have used this value in our calculations. The contribution of core electrons has been described by norm-conserving Troullier–Martins relativistic pseudo-potentials,³³ factorized in the Kleinman–Bylander form.³⁴ Valence electrons were treated explicitly using a split-valence basis set of double- ξ plus polarization functions obtained with an energy shift of 10 meV.³⁵ The energy cutoff for the real space integration mesh was set to 300 Ry and the Brillouin zone was sampled using a mesh with $9 \times 25 \times 3$ k-points obtained using the method of Monkhorst and Pack.³⁶ All calculations were performed using the experimental crystal structure.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This research was supported by the French National Research Agency Grants, ANR 16-ACHN-0007 and ANR 17-ERC3-0003, and a PhD grant (to M. Beau) from the Bretagne region. Y. Kiyota is grateful to JSPS for a research fellowship (17J05098). We thank Dr. A. Vacher (Rennes) for the preparation of (*Z,E*)-(SCH₂CH₂CN)₂(Me)₂TTF and CDIFX-Rennes for access to X-ray diffractometer. Work in Spain was supported by MICIU (PGC2018-096955-B-C44 and PGC2018-093863-B-C22), MINECO through the Severo Ochoa (SEV-2015-0496) and Maria de Maeztu (MDM-2017-0767) Programs and Generalitat de Catalunya (2017SGR1506 and 2017SGR1289). Work in Japan was supported by JPSJ KAKENHI (Grant Number 16K13974 and 18H02044) and Takahashi Industrial and Economic Research Foundation.

Notes and references

‡ Crystal data for [(*E*)-1](TCNQ): C₂₂H₁₆N₄S₆, $M = 528.75$, triclinic, space group $P-1$, $T = 150(2)$ K, $a = 3.8415(3)$ Å, $b = 7.6270(7)$ Å, $c = 18.9474(17)$ Å, $\alpha = 90.790(3)^\circ$, $\beta = 90.292(3)^\circ$, $\gamma = 93.350(3)^\circ$, $V = 554.13(8)$ Å³, $\rho_{\text{calcd}} = 1.584$ g cm⁻³, $R_1(wR_2) = 0.0377(0.0905)$ and $S = 1.133$ for 2519 reflections and 2217 reflections with $I > 2\sigma(I)$. CCDC 1901937.

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