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Esterification of Levulinic Acid with Butanol over ion exchange resins.

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The authors regrets that there are amendments in Table 4. The correct version is shown below.

Table 4. NH₃ adsorption enthalpy and acid capacity from NH₃ adsorption calorimetry.

Catalyst	$-\Delta H_{\text{NH}_3}$ (kJ/mol)	[H ⁺] (mmol/g) ^a	[H ⁺] (mmol/g) ^b	Ref
Amberlyst 15	111 ± 2	4.70 ± 0.1	4.81	48
	110 ± 3	4.73 ± 0.1		80
Amberlyst 35	117 ± 2	5.20 ± 0.1	5.32	48
Amberlyst 16	108 ± 3	4.72	4.80	80
Amberlyst 36	117 ± 2	5.30 ± 0.1	5.40	48
Amberlyst 39	111 ± 3	4.66	4.82	80
Amberlyst 70	117 ± 3	1.65 ± 0.05	2.55	48
Dowex 50Wx4	113 ± 3	4.65	4.95	80
Dowex 50Wx2	106 ± 3	3.80	4.83	80
Purolite CT224	112 ± 3	4.51	5.34	80
Amberlyst 46	108 ± 3	0.91	0.87	80

The experimental technique can be found in Siril et al. [48]

^a From microcalorimetry of NH₃ adsorption. It considers those centers with $-\Delta H_{\text{NH}_3} \geq 80$ kJ/mol

^b From titration with NaOH

References cited

[48] P.F. Siril, H.E. Cross, D.R. Brown, *J. Mol. Catal. A: Chem.* 279 (2008) 63–68.

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The authors would like to apologize for any inconvenience caused

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